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10.1 Equation Stoichiometry 371 There is a shortcut for this calculation. We can collapse all five of the conversion factors above into one. The reaction equation tells us that there are six moles of  $H_2O$  for each mole of  $P_4O_{10}$

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10.1 Equation Stoichiometry Section 6.7 asked you to pretend to be an industrial chemist at a company that makes phosphoric acid,  $H_3PO_4$ . The three-step “furnace method” for producing this compound is summarized by these three equations: ...

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© 2011 Pearson Education, Inc. Chapter 10 2 What Is Stoichiometry? • Chemists and chemical engineers must perform calculations based on balanced chemical reactions to predict the cost of processes. • These calculations are used to avoid using large, excess amounts of costly chemicals. • The calculations these scientists use are called

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Example 10.1 shows how the coefficients in a balanced chemical equation provide a number of conversion factors that allow us to convert from moles of any reactant or product to moles of any other reactant or product. example 10.1 - Equation Stoichiometry

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Chapter 3: Stoichiometry: Chemical Calculations: ... In some stoichiometry problems the limiting reactant must first be identified through a preliminary calculation. 3.10 Yields of Chemical Reactions—The calculated quantity of a product is the theoretical yield of a reaction.

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