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(ii) CO_2 (iii) CH_4 Answer: (i) Molecular

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mass of $H_2O = 2(1.008 \text{ amu}) + 16.00 \text{ amu} = 18.016 \text{ amu}$ (ii) Molecular mass of $CO_2 = 12.01 \text{ amu} + 2 \times 16.00 \text{ amu} = 44.01 \text{ amu}$ (iii) Molecular mass of $CH_4 = 12.01 \text{ amu} + 4(1.008 \dots$

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Papers. Question 1. (i) Calculate the number of electrons which will together weigh one gram. (ii) Calculate the mass and charge of one mole of electrons.

Answer: (i) Mass of an electron = 9.1×10^{-28} g.

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$K.[N_2(g)] = 1.5 \times 10^{-2} M$ $[H_2(g)] = 3.0 \times 10^{-2} M$ $[NH_3] = 1.2 \times 10^{-2} M$.
Calculate equilibrium constant.

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