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Modern Chemistry 2 Measurements and Calculations CHAPTER 2 REVIEW Measurements and Calculations SECTION 1 SHORT ANSWER Answer the following questions in the space provided.

1. Determine whether each of the following is an example of observation and data, a theory, a hypothesis, a control, or a model.

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2 Measurements and Calculations

Chapter 2: Measurements and Calculations There is a lot of math involved in the study of chemistry. Therefore, it is important for to review the metric system and some of the units we will be using...

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SECTION 2 continued 4. a. Write the equation for the first

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ionization of H_2CO_3 in aqueous solution. Assume that water serves as the reactant that attaches to the hydrogen ion released from the H_2CO_3 . Which of the reactants is the Brønsted-Lowry acid, and which is the Brønsted-Lowry base? Explain your answer. $\text{H}_2\text{CO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{HCO}_3^- + \text{H}_3\text{O}^+$

14 Acids and Bases

CHAPTER 11 REVIEW Gases Class SHORT ANSWER Answer the following questions in the space provided. Increasing rate of effusion. List the following gases in order (Assume all gases are at the same temperature and pressure.) (c) HCl (d) O_2 . Explain your reasoning for the order of gases you chose in item 1 above. Refer to the kinetic-

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CHAPTER 11 REVIEW Gases SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. State whether the pressure of a fixed mass of gas will increase, decrease, or stay the same in the following circumstances: a. temperature increases, volume stays the same b. volume increases, temperature stays the same

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$2(\text{SO}_4)^{2-} + 3(\text{s}) \rightarrow 2\text{Fe}_3(\text{aq}) + 3\text{SO}_4^{2-}(\text{aq})$ 5. a. Write the net ionic

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equation for the reaction that occurs when solutions of lead(II) nitrate and ammonium sulfate are combined. $\text{Pb}^{2+}(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) \rightarrow \text{PbSO}_4(\text{s})$ b. What are the spectator ions in this system? NO_3^- and NH_4^+ are spectator ions. MODERN CHEMISTRY IONS IN AQUEOUS SOLUTIONS AND ...

13 Ions in Aqueous Solutions and Colligative Properties

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