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9.4: Properties of Solutions Vapor Pressure Depression. All liquids evaporate. In fact, given enough volume, a liquid will turn completely into a... Boiling Point and Freezing Point Effects.

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A related property of solutions is that their boiling points are higher than... Osmotic Pressure. The last ...

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between dilute solution
and pure solvent
depends on the
amount of solute
present not its identity
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problems (from pages
28-29): 1. What is the
mass percent of NaCl
when 5.00 g of NaCl is
dissolved in 95.0 g of
water? Answer: 5.00 %

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by mass NaCl Set-up:
100 5.00 95.0 5.00 100
solution mass solute
mass % x gNaCl gwater
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formation of solution
can be either
exothermic or

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endothermic;
exothermic processes
are spontaneous;
solution will not form if
enthalpy is too
endothermic; H_3 has
to be comparable to $H_1 + H_2$. Ionic
substances cannot
dissolve in nonpolar
liquids; Polar liquids do
not form solutions with
nonpolar liquids; 13.1.2
Solution Formation,
Spontaneity, and
Disorder

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13.S: Properties of Solutions (Summary) **- Chemistry** **LibreTexts**

Chapter 16.1 //

Properties of Solutions.

Saturated solution.

Solubility. Unsaturated solution. Miscible. A solution containing the maximum amount of solute for a given....

The amount of a substance that dissolves in a given quantity o.... A solution that contains less

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solute than a saturated solution...

ch. 16.1 properties of solutions Flashcards and Study Sets ...

33. Solution A, it has the greatest concentration of ions.

34. 500. mL of 4.0M NaOH contains 2 mol NaOH, which is 80. g NaOH. Place 80. g NaOH in a graduated container. Add enough distilled water to

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dissolve the NaOH by swirling or stirring. When the solute is dissolved, add enough distilled water to make 500. mL of solution.

ANSWERS TOPIC 7

Part C Review

Questions 1. 4 2. 2

3. 3 1 ...

The properties of solutions are very similar to the properties of their respective pure solvents. This makes

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sense because the majority of the solution is the solvent.

However, some of the properties of solutions differ from pure solvents in measurable and predictable ways.

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SECTION 16.1

PROPERTIES OF

SOLUTIONS (pages

471-477) This section

identifies the factors

that affect the

solubility of a

substance and

determine the rate at

which a solute

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dissolves. Solution Formation (pages 471-472) Look at Figure 16.1 on page 471 to help you answer Questions 1 and 2.

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2.3: Chemical Properties of Carboxylic Acids-

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Choose an answer and hit 'next'. You will receive your score and answers at the end.
question 1 of 3. ... To learn more about the properties of solutions, review the corresponding lesson called ...

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- Notes 12.3

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12.3 Concentrations
Answer Key - Practice
Solution Equations (link
will be posted at
9:15am in google
classroom) - Worksheet
12.3 Concentrations
(link will be posted by
10:00am in google
classroom)

Unit 4.2: Solutions and Colligative Properties (modified

...

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