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## **The Nature Of Gases Section**

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## **CHEMISTRY: CHAPTER 13 SECTION 1: THE NATURE OF GASES ...**

Nature of Gases: We are constantly surrounded and buried in gas molecules. to move freely through them while constantly filling in behind us. We breath them in and out and use them to make energy within our bodies. They

## **The Nature of Gases - California State University, Northridge**

A gas is composed of particles molecules or atoms Considered to be hard spheres

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far enough apart that we can ignore their volume. Between the molecules is empty space. The Kinetic Theory of Gases Makes three descriptions of gas particles

## **Chapter 13- The States of Matter**

### **13.1- The Nature of Gases**

SECTION 13.1 THE NATURE OF GASES (pages 385–389) This section introduces the kinetic theory and describes how it applies to gases. It defines gas pressure and explains how temperature is related to the kinetic energy of the particles of a substance. Kinetic Theory and a Model for Gases (pages 385–386) 1.

### **Name Date Class STATES OF MATTER 13**

considered to be small, hard spheres with an insignificant volume. Particles are relatively far apart compared to solids and liquids. Between particles there is empty space. No attractive or repulsive forces exist between the particles. The motion of one particle in a

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gas is independent of the motion of all the other particles.

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Gas, one of the three fundamental states of matter, with distinctly different properties from the liquid and solid states. The remarkable feature of gases is that they appear to have no structure at all. They have neither a definite size nor shape, whereas ordinary solids have both a definite size

## **Gas | state of matter | Britannica**

The nature of gases 131. Diffusion elastic kinetic molecular theory. Results from the force exerted by a gas per unit surface area of an object. Because during an evaporation particles with highest kinetic energy tends to escape first and the particles left have lower kinetic energy making it cool.

## **131 The Nature Of Gases Worksheet - Blogger**

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## **nature of gases section review answer key**

kinetic theory. According to the kinetic theory, all matter consists of tiny particles that are in constant motion. Theory #1. The particles in a gas are considered to be small, hard sphere with an insignificant volume. Theory #2. The motion of the particles in a gas is rapid,

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constant and random. Theory #3.

## **Chemistry 13.1 The Nature of Gases Flashcards | Quizlet**

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Properties of Gases The ideal gas model is used to predict changes in four related gas properties: volume, number of particles, temperature, and pressure. Volumes of gases are usually described in liters, L, or cubic meters, m<sup>3</sup>, and numbers of particles are usually described in moles, mol.

## **Chapter 13 Gases - An Introduction to Chemistry**

SECTION 13.1 THE NATURE OF GASES 1. Explain why there is no gas pressure inside a vacuum. 2. How would the reading on a barometer change if you

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were to take one on a trip from Los Angeles to Lake Tahoe, which is at a much higher altitude? 3. The height of a column of mercury in a barometer is 754.3 mm. What is the atmospheric

## **Chem ID# Name CHAPTER 13 STATE OF MATTER SECTION 13.1 THE ...**

Gases are easily 1 because of the 2 between particles in a gas. The four variables used to describe a gas are pressure, (P), 3 4 2 S 5. 8. dec a g h5 No are c n and number of (n). You can use 6 theory to predict and explain how gases will respond to a change in conditions. Doubling the amount of gas in a rigid container 7 the pressure. You can raise the

## **Bozeman Public Schools**

The Nature of Gases>Kinetic Theory and a Model for Gases. According to kinetic theory: •The particles in a gas are considered to be small, hard spheres with an insignificant volume. •The motion of the particles in a gas is rapid,

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constant, and random. •All collisions between particles in a gas are perfectly elastic.

## **What do I already know about states of matter? (index card ...**

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Section: The Nature of Science 1. Name four branches of biological science. 2. ... Identify each of the following as a gas, liquid, solid, or plasma. \_\_\_\_ a. The particles are closely packed together, ... Section: Behavior of Gases 1. Identify which gas law is being demonstrated.

## **Physical Science Concept Review Worksheets with Answer Keys**

L. Anderson



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## **L. Anderson**

gases are similar, and one way in which they are different. Both liquids and gases can flow, so they can take the shape of their container. The molecules in a liquid have intermolecular attractions that are not present in gases. Therefore, liquids have a definite volume and will not simply fill their container.

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